M.PHIL (CHEMISTRY : SEM -I) PAPER III – [Elective] PHYSICAL CHEMISTRY [50 MARKS]

Corrosion (25Marks)

[25 Hours]

- 1. Basic principles of corrosion thermodynamics and electrode kinetic considerations acidic, basic and neutral solutions corrosion by oxygen corrosion by soils Aerobic and anaerobic micro organisms oxidation of metals and alloys mechanism of various growth laws study of oxide films.
- Corrosion by acids oxidizing and non oxidizing acids hydrogen cracking.
 Atomosphetic corrosion Modern mechanism, passivity of metal alloys. Relative merits of material of construction for industries and their selection considerations.
 Corrosion characteristics of non ferrous metals and alloys stainless steels.

References:

- 1. An introduction to metallic corrosion --By Ulick R. Evans
- 2. Corrosion Corrosion f metals and alloys–Corrosion control Vol: 1&2

--By L.L.Shreir {George Newness Ltd.}

2. Electrochemistry {25 marks }

[20 Hours]

- The theory of electrolytic dissociation the process of electrolytic dissociation and conductance Electrolytic conductance of solutions the migration of ions Electrokinetic phenomena structure and properties of double layer electro osmosis The rate of electro osmotic flow the pressure equation of Electro electro stenolysis.
- Electromotive force Electro motive force of concentration cells the potential difference at the liquid liquid function flowing functions. Gas electrodes Applications of potentiometric measurements EMF of oxidation reduction cells polarization and electrolysis Decomposition voltage over voltage Electrolytic separation of metals polarographic analysis The Edision accumulator De polarization and electrolysis.

References:

- 1. Principles and Applications of Electrochemistry
 - -- By H.Jermain Creilghton. {John Wiley & Sons }

- 2. Comprehensive Treatise of Electrochemistry
 - -- Vol.:1&2-By J.O.M.Bockris,Brian. F. Conway & Ernest Yeager
- 3. Modern aspects of electrochemistry
 - -- By B.E.Conway and J.O'M.Bockris [Plenum press]